

2nd Year

Chemistry Section-I

Marks: 30

T1: Ch. # 1

Time: 60 Mints

Question No. 1: Encircle the most suitable answer.

(6x1=6)

1. Which of the following is intermediate hydride?

- a) NaH b) BeH₂ c) CH₄ d) CaH₂

2. Which of the following is more ionic?

- a) MgCl₂ b) AlCl₃ c) SiCl₄ d) PCl₅

3. The most acidic nature is shown by:

- a) Mn₂O b) Mn₂O₃ c) MnO₂ d) MnO

4. The oxides of electropositive elements are:

- a) Basic b) Acidic c) Amphoteric d) Neutral

5. Keeping in view the size of atoms, which one order is correct?

- a) Mg > Sr b) Ba > Mg c) Lu > Ce d) Cl > I

6. The metal which is used in thermite process because of activity:

- a) Iron b) Copper c) Aluminum d) Zinc

Section-II

Note: Give short answers of any 8 Questions.

(8x2=16)

1. How do you justify that ZnO is an amphoteric oxide?
2. What is shielding effect? Give example.
3. Describe the trends of melting and boiling points from left to right in a period.
4. How hydrogen resembles with IVA group elements?
5. Why noble gases have zero oxidation state?
6. What are the improvements made in Mendeleev's periodic table?
7. Why the ionic sizes of negative ions are larger than their parent atoms?
8. Why metallic character of metals increases from top to bottom in a group of metals?

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Section-III

(a) How can you justify the position of Hydrogen in group I-A? Give similarities and dissimilarities? (4)

(b) What is modern periodic table also discuss hydration energy in detail. (4)

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